

2-10-2015

A Simple Method for Tuning the Glass Transition Process In Inorganic Phosphate Glasses

René Fulchiron
Université de Lyon

Imane Belyamani
University of Southern Mississippi

Joshua U. Otaigbe
University of Southern Mississippi

Véronique Bounor-Legaré
Université de Lyon

Follow this and additional works at: https://aquila.usm.edu/fac_pubs

 Part of the [Chemistry Commons](#)

Recommended Citation

Fulchiron, R., Belyamani, I., Otaigbe, J. U., Bounor-Legaré, V. (2015). A Simple Method for Tuning the Glass Transition Process In Inorganic Phosphate Glasses. *Scientific Reports*, 5, 1-5.
Available at: https://aquila.usm.edu/fac_pubs/16693

This Article is brought to you for free and open access by The Aquila Digital Community. It has been accepted for inclusion in Faculty Publications by an authorized administrator of The Aquila Digital Community. For more information, please contact Joshua.Cromwell@usm.edu.



OPEN

SUBJECT AREAS:

CHEMISTRY

MATERIALS SCIENCE

OPTICS AND PHOTONICS

A simple method for tuning the glass transition process in inorganic phosphate glasses

René Fulchiron^{1,2,3}, Imane Belyamani⁴, Joshua U. Otaigbe^{1,2,3,4*} & Véronique Bounor-Legaré^{1,2,3}Received
2 October 2014Accepted
9 January 2015Published
10 February 2015Correspondence and
requests for materials
should be addressed to
J.U.O. (Joshua.
Otaigbe@usm.edu)* Current address:
School of Polymers and
High performance
Materials, The
University of Southern
Mississippi,
Hattiesburg,
Mississippi 39406,
USA.

¹Université de Lyon, F-69361 Lyon, France, ²CNRS, UMR 5223, Ingénierie des Matériaux Polymères, F-69621 Villeurbanne, France, ³Université Claude Bernard Lyon 1, F-69622 Villeurbanne, France, ⁴School of Polymers and High performance Materials, The University of Southern Mississippi, Hattiesburg, Mississippi 39406, USA.

The physical modification of glass transition temperature (T_g) and properties of materials via blending is a common practice in industry and academia and has a large economic advantage. In this context, simple production of hitherto unattainable new inorganic glass blends from already existing glass compositions via blending raises much hope with the potential to provide new glasses with new and improved properties, that cannot be achieved with classical glass synthesis, for a plethora of applications such as computers screens, glass-to-metal seals, and storage materials for nuclear wastes. Here, we demonstrate that blends of the specific glass compositions studied are miscible in all proportions, an unreported phenomenon in hard condensed matter like glass. Interestingly, excellent agreement was found between the obtained data and calculated T_g s from theoretical equations (Supplementary information) for predicting the composition dependence of T_g for miscible blends with weak but significant specific interactions between the blend components. That this blending method is at present not applied to inorganic glasses reflects the fact that water and chemically resistant phosphate glasses with relatively low T_g s have become available only recently.

In general, when a liquid phase is cooled to below its freezing temperature, it normally transforms into a crystalline solid as Fig. 1 shows^{1–7}. However, inorganic glass liquids like phosphate glasses do not crystallize but instead form a rigid disordered network structure when cooled to low temperatures due to their complex molecular configuration or relatively slow transport. The disordered structure just mentioned is remarkably similar to that of the liquid phase. In contrast, for polymeric liquids, crystallization is complicated because of their monomer chain lengths and weak Van der Waals forces that permit individual chain molecules to readily slide past one another^{8,9}. It is the complex structural rearrangements required for crystallization that leads to glass formation in both glasses and polymers. It is now generally recognized that inorganic phosphate glasses are polymeric in nature in the sense that they are composed of chain-like or crosslinked structures like organic polymers, albeit the chains (composed of phosphate anion tetrahedra) are much shorter in the former case^{10–17}.

It is worthy to note that the glass transition temperature (or T_g), which corresponds to the temperature at which a supercooled liquid becomes a glass, is for inorganic glasses what the melting point (or T_m) is for crystalline solids (see Fig. 1). While T_m values do not depend on the direction of the change (i.e., freezing a liquid, melting a solid) or on the change rate, the location of the glass transition region depends on both factors as depicted in Fig. 1. Kovacs showed in 1963 in his classical paper how the T_g location varies with the cooling rate of ‘glassy’ polyvinyl acetate¹⁸.

It is well known that T_g values are useful for a variety of purposes, especially needed are T_g values as a function of composition for binary polymer blends because they tell us whether the blend components are miscible, or compatible, or not miscible at all. In this context, a single glass transition temperature for all the blends characterizes complete miscibility. On the other hand, compatible systems show two T_g values^{19,20} which depend on composition while useful immiscible polymers show that the T_g values for the pure components do not change with composition. Note that the miscibility (or lack of it) is often decisive for all properties for both inorganic glasses and polymers including, for polymers, effects of fillers, nano-confinement effects on segmental motions in polymer composites, and changes of T_g with residual stress^{21,22}.

While T_g as a function of composition has been used widely to develop useful polymer blends and copolymers, and other simple glassy liquids and thin films from aqueous emulsions and paints, surprisingly it has not been used to develop inorganic glasses with hitherto unattainable T_g values and properties from existing compositions considering the facile and relative technological importance of this approach. The blends behavior and the

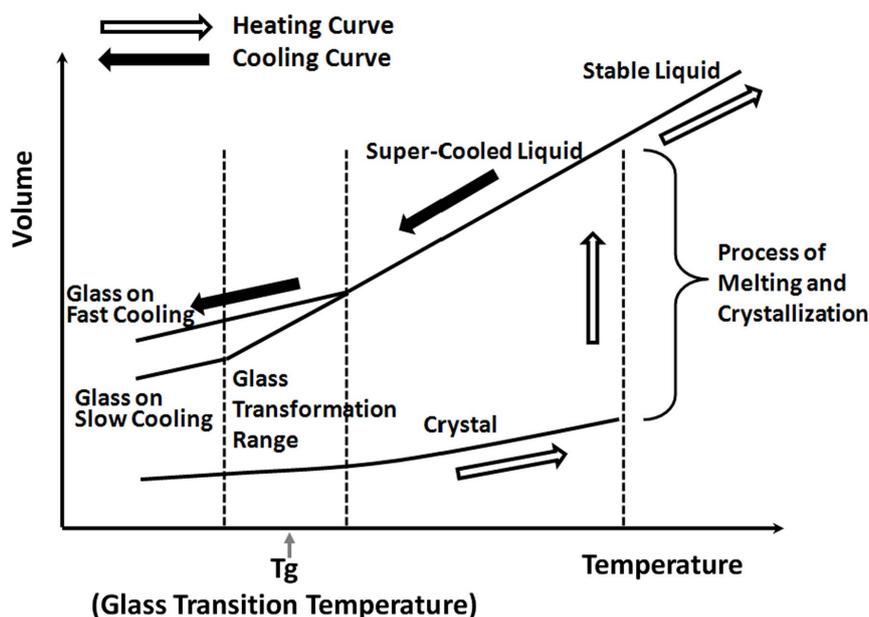


Figure 1 | General volume changes associated with heating and cooling in systems susceptible to glass formation (Adapted from Refs. 1–7).

variation of T_g of silicate glasses with composition have not been systematically studied, probably because of the difficulties of blending them at elevated temperatures (often $>1000^\circ\text{C}$) in the liquid state and of accurately measuring T_g in the very high temperature ranges typical of conventional inorganic glasses. The unique desirable possibility of tailoring T_g of inorganic phosphate glasses via blending should extend the versatile, low cost, and facile blending method to a wide variety of inorganic glass compositions that may lead to new applications such as optoelectronic and biocompatible biomedical devices, and storage materials for nuclear wastes²³ where already existing glasses cannot be used, essentially similar to current practice in the polymer industry.

In this paper we report the first ever successful attempt of blending 'premade' (or existing) phosphate glass compositions in the liquid phase to obtain new, hitherto unattainable miscible compositions with a different single T_g that varies with the blend composition. The obtained results are analyzed and interpreted within the context of a number of extant theoretical equations that pervade the literature for predicting the glass transition temperatures of binary blends and copolymers (see Supplementary information). Like for organic polymers, the properties of the phosphate glass blends can be expected to be some combination of that of the blend components to a greater or lesser extent depending on the exact compositions. This is a matter for future investigation.

In the current study described in this paper, representative existing low T_g tin-phosphorous oxyfluoride and mixed alkali glass systems were used as good model systems because of their ease of preparation and handling, durability to water and chemicals, low working temperatures (low- T_g s compared to other commonly used inorganic glasses), and sensitivity of intermediate-range order to melt-processing conditions^{10,24}. It is hoped that the interesting results of this study will provide a basis for further exploration of the facile idea of developing new phosphate glass compositions from already existing glass compositions.

Figure 2a–b shows the DSC traces of the second heating for the pure starting phosphate glass (P-glass) compositions and of the first heating for the blend sample LTw0.56, respectively. Two T_g values are clearly noticeable, each of them being due to the solid-state powder of initial components in the blend. After this first heating and the subsequent cooling, the second heating of this blend (Fig. 2c) still shows two T_g values, revealing that the blend is not totally achieved.

However, the first T_g is much more discernible than the second one. Additionally, each T_g moved to a different location compared to that shown in Fig. 2a because the P-glass blends are partially mixed and there are two phases with one phase richer in ILT2 ($T_g = 178^\circ\text{C}$) and the other phase richer in IHT2 ($T_g = 276^\circ\text{C}$). For the third heating, practically only one T_g (186°C) was observed (see Fig. 2d). Therefore, it can be concluded that the total intimate mixing of the P-glass blend is nearly achieved. However, it must be recognized that Fig. 2d still shows a small thermal event around 275°C , probably due to a very small part of the remaining phase that is richer in IHT2. Nevertheless, after another holding at 370°C for 2 hours (not shown here) the heating curve does not exhibit this small event anymore and there is only one T_g at practically the same value as shown here ($\sim 186^\circ\text{C}$).

Now considering the other blends in Table 1, Fig. 2e shows the DSC trace obtained for the P-glass blend sample LTw0.24, after holding at 370°C for 2 hours. Clearly this figure shows that only one T_g (245°C) is discernible, indicating again that a completely miscible P-glass blend was achieved. Additionally, Fig. 2f shows the DSC curve of the third heating of the P-glass blend sample LTw0.78 illustrating total miscibility of the blend components as depicted by the unique single glass transition temperature indicated.

Finally, all the T_g values obtained (taken at midpoints of the glass transition region) are plotted as a function of the IHT2 weight fraction in Fig. 3. From this Figure, it can be concluded that the classical Fox Equation (Supplementary equation (1)) fails in describing the T_g variation of the blends of the phosphate glass compositions studied. From the theoretical point of view, this finding leads to the conclusion that specific interactions between the P-glass blend components (ILT2 and IHT2), favorable to their miscibility, are present in the final P-glass blends. Admittedly, the experimental variation shows a slight negative deviation from the simple mixing rule but much lesser than that predicted by the Fox equation which is strictly valid for blends without any specific interactions (Supplementary information). Interestingly, the experimental data trend is well described by the Gordon-Taylor equation (Supplementary equation (2)). The adjustable parameter K , obtained by fitting the experimental results to Supplementary equation (2) was found to be 0.883. By analogy with the expression of Lu and Weiss, (Supplementary equations (17), (18), and (14)), this obtained value of K which is quite close to 1 leads to the conclusion that some relatively weak but significant specific

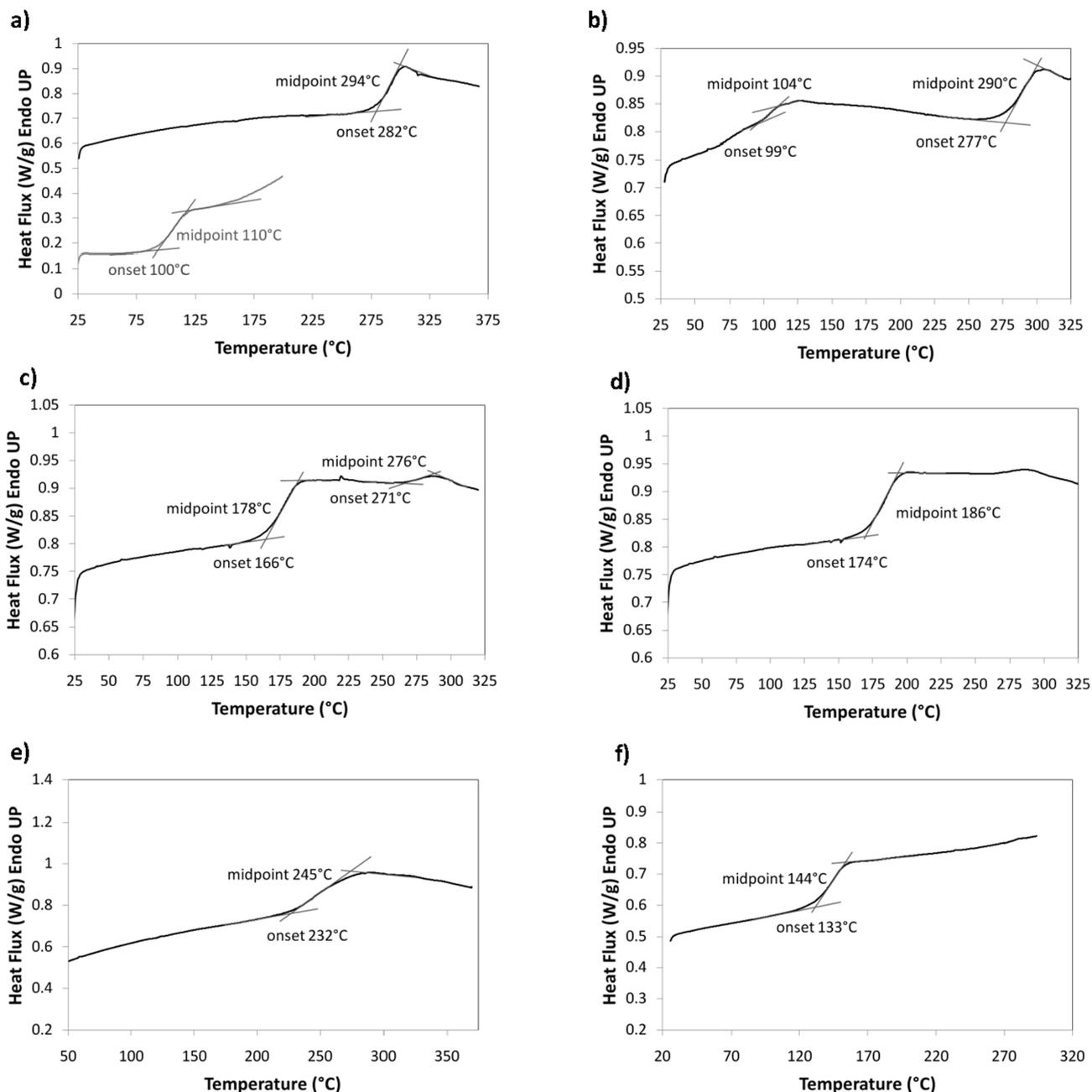


Figure 2 | DSC traces: (a) Second heating for starting compositions ILT2 (bottom) and IHT2 (top), (b) first heating for LTw0.56, (c) second heating for LTw0.56, (d) third heating for LTw0.56, (e) second heating after holding at 370 °C for 2 hours for LTw0.24, (f) third heating for LTw0.78.

interactions are present between both components of the P-glass blends studied. The specific interactions just mentioned are ascribed to the well known ionic interactions and liquid state re-ordering of phosphate anion tetrahedral chains present in phosphate glasses as

previously reported in the literature by Otaigbe and Beall¹⁰ and by other glass science researchers²⁵. It is noteworthy that strong interactions (if present) would instead lead to a positive deviation from the mixing rule.

Table 1 | Samples composition and temperature profiles used

Sample name	ILT2	LTw0.78	LTw0.56	LTw0.24	IHT2
Weight fraction of ILT2	1.000	0.783	0.558	0.244	0.000
Weight fraction of IHT2	0.000	0.217	0.442	0.756	1.000
T_{max} (°C)	200	320	370	370	370
Holding time	2 min	2 min	2 min	2 h	2 min
Number of cycles	2	3	3	2	2

All heating and cooling rates are 30 °C/min.

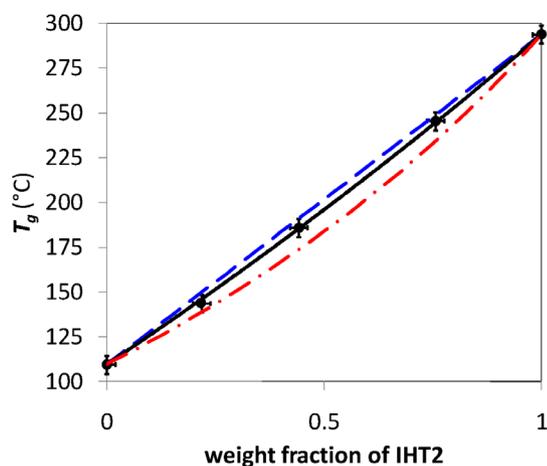


Figure 3 | T_g versus weight fraction of IHT2. ●: Experimental values, ---: Linear mixing rule, — · —: Fox equation (Supplementary equation (1)), — Gordon Taylor equation (Supplementary equation (2) with $K = 0.883$). The error bars shown are ± 0.01 and $\pm 2^\circ\text{C}$ for weight fraction and T_g , respectively.

In conclusion, this study unambiguously demonstrates that phosphate glass compositions with significantly different T_g values (and therefore properties) can be easily blended under controlled heating conditions in the liquid phase to obtain new phosphate glass compositions with a different single T_g and properties. Interestingly, the obtained results show that the specific phosphate glass compositions used in this study are miscible in all proportions as predicted by a number of empirical and quantitative equations in the literature for interpreting the composition dependence of T_g for miscible blends with relatively weak but significant specific interactions already described (see Supplementary information). It is noteworthy that the observed phenomenon just mentioned is hitherto unreported for hard condensed matter like inorganic glasses. While the facile blending strategy reported here is used widely for organic polymer blends and copolymers, to our knowledge, this is the first published experimental study of application of this strategy to inorganic glasses to obtain new industrially useful inorganic glasses with enhanced benefits from already existing glass compositions, providing a distinctive transformative and complementary alternative method to the traditional relatively cumbersome chemical synthesis of new glasses with new T_g and properties. With additional engagement of the scientific and industrial communities, it is hoped that the current research finding can be accelerated into glass products with enhanced benefits that includes conservation of naturally occurring oxide compounds used in conventional glass synthesis. Conceptually, it may even be possible to use this simple practical approach to tune important glass application properties like chemical durability of phosphate glasses in general, in a manner that is consistent with Einstein's quote, "Everything should be made as simple as possible, but not simpler"²⁶. Thus, making the strategy potentially widely applicable. The current paper will guide and facilitate future progress in this important and emerging area.

Methods

Materials. A comprehensive literature review revealed a number of patents illustrating the preparation of inorganic glasses with different T_g values based on inorganic oxides and phosphates. The T_g values are related to the type and the stoichiometric molar concentrations of the reactants used. In this study, phosphate glass (P-glass) compositions with two remarkably different glass transition temperatures and compositions were synthesized in the laboratory according to procedures reported elsewhere by Otaigbe and Beall¹⁰, Sanford and Tick²⁷, Beall and Quinn²⁸, and Beall and Pierson²⁹. Typical compositions of tin fluorophosphate glass (hereinafter denoted as ILT2) and of a mixed alkali phosphate glass (hereinafter denoted as IHT2) compositions were prepared by conventional glass melting and quenching method. In the current study, the ILT2 sample with a molar composition

of 10% SnO + 45% SnF₂ + 42.8% P₂O₅ (all expressed as mol%, and optionally containing 2.2% tin pyrophosphate additive) was prepared by heating, in a furnace maintained at 420°C, a mixture of the stoichiometric amounts of the initial raw materials in an appropriate high-temperature crucible for at least 25 minutes. The glass melt was then poured into a steel mold to form a small circular disk that was subsequently annealed at a temperature of 90°C for 2 hours. The obtained glass was then ground into powder to give the ILT2 sample with a T_g of approximately 110°C. A similar procedure was used to prepare the IHT2 sample of this study with a molar composition of 50% M₂O + 2% Al₂O₃ + 48% P₂O₅ (where M is an alkali such as Na or Li, and all expressed in mol%). The stoichiometric amounts of the initial raw materials of the IHT2 sample were heated at 1000°C for at least 2 hours, and the resulting glass melt was poured into a steel mold to form a circular disk that was subsequently annealed at 260°C for 2 h to give a relatively high T_g of ~294°C. It is worthy to note that these glasses and other similar glasses can be easily prepared and handled as previously reported^{10,24}.

Samples preparation. The P-glass samples, in powder form, were simply mixed by weighing the required amounts of each material in a vial and stirring vigorously by hand. The different weight fractions of the used samples are given in Table 1:

Thermal Analysis. The thermal properties of the samples were acquired using a Differential Scanning Calorimetry Diamond DSC from Perkin Elmer using between 20 and 30 mg of P-glass blend sample in a nitrogen atmosphere following procedures reported in the literature³⁰. For the initial starting phosphate glasses (i.e., samples ILT2 and IHT2), two cycles of heating and cooling were applied from ambient temperature to the maximum temperature (T_{max}) at heating and cooling rates of 30°C/min. These rates are higher than that frequently applied (10°C/min) in standard DSC measurements in the literature because a higher rate is generally suitable for T_g analysis due to the experimental fact that the heat flux signal is enhanced so that the variation of the heat capacity characteristic of the T_g transition is more easily detected. However, these rates must not be too high otherwise the accuracy of the temperature scale can be reduced. Between the heating and the cooling, the sample was held at T_{max} for 2 minutes. The maximum temperature used was: 200°C for ILT2 and 370°C for IHT2. In fact, for ILT2, when the temperature was increased to more than 200°C, some instabilities of the signal appeared, due to boiling of the sample in the DSC sample pan. For the blend samples, the imposed temperature profiles were slightly different because the intimate mixing of the materials in the DSC sample pan has to be ensured which means that the temperature T_{max} must be higher than the T_g of both materials, but T_{max} and the holding time must not be too high to avoid possible desegregation of the ILT2 material from the blend (especially for the highest content of ILT2). Note that the T_{max} , the holding time at T_{max} , and the number of heating and cooling cycles applied to the samples in Table 1 were found to give accurate and reproducible results.

Extant theoretical considerations. Details of the theoretical considerations are given in the Supplementary Information.

- Kingery, W. D., Bowen, H. C. & Uhlmann, D. R. *Introduction to Ceramics (2nd edn.)* (Wiley and Sons, New York, 1976).
- Guy, A. J. *Essentials of Materials Science* (McGraw-Hill, New York, 1976).
- Varshneya, A. K. *Fundamentals of Inorganic Polymers* (Academic Press, North-Holland, 1994).
- Paul, A. *Chemistry of Glasses (2nd edn.)* [17–29] (Chapman & Hall, London, 1990).
- Vogl, W. *Chemistry of Glass (2nd edn.)* (The American Ceramic Society, Ohio, 1985).
- Debenedetti, P. G. & Stillinger, F. H. Supercooled liquids and the glass transition. *Nature* **410**, 259–267 (2001).
- Angell, C. A., Sare, J. M. & Sare, E. J. Glass transition temperatures for simple molecular liquids and their binary solutions. *J. Phys. Chem.* **82**, 2622–2629 (1978).
- Wunderlich, B. *Macromolecular Physics Vol. 1 Crystal structure, Morphology, Defects* (Academic Press, North-Holland, 1976).
- Piorkowska, E. & Rutledge, G. C. (eds.). *Handbook of Polymer Crystallization* (Wiley, New York, 2013).
- Otaigbe, J. U. & Beall, G. H. Inorganic phosphate glasses as polymers. *Trends Polym. Sci.* **5**, 369–379 (1997).
- Van Wazer, J. R. in *Inorganic Polymers* [Stone, F. G. A. & Graham, W. A. G. (eds.)] (Academic Press, North-Holland, 1962).
- Eisenberg, A. Ionic forces in polymers: Some properties of the polyphosphates in the glass transition region. *Adv. Polym. Sci.* **5**, 59–112 (1967).
- Ray, N. H. The structure and properties of inorganic polymeric phosphates. *Br. Polym. J.* **11**, 163–177 (1979).
- Abe, Y. in *Topics in Phosphorous Chemistry* [Vol. 11] [Grayson, M. & Griffith, E. J. (eds.)] [19–67] (Wiley, New York, 1983).
- Mark, J. E., Allcock, H. R. & West, R. *Inorganic Polymers* (Prentice Hall, New Jersey, 1992).
- Warren, B. E. Fourier analysis of x-ray patterns of soda-silica glass. *J. Am. Ceram. Soc.* **1938**, **21**, 259–265 (1938).
- Zachariasen, W. H. The atomic arrangement in glass. *J. Am. Chem. Soc.* **54**, 384–351 (1932).



18. Kovacs, A. J. Glass transition in amorphous polymers. Phenomenological study. *Adv. Polym. Sci.* **3**, 394–508 (1963).
19. Lodge, T. P. & McLeish, T. C. B. Self-Concentrations and Effective Glass Transition Temperatures in Polymer Blends. *Macromolecules*. **33**, 5278–5284 (2000).
20. Lodge, T. P., Wood, E. R. & Haley, J. C. Two calorimetric glass transitions do not necessarily indicate immiscibility: the case of PEO/PMMA. *J. Polym. Sci. Part B: Polym. Phys.* **44**, 756–763 (2006).
21. Thomas, S. & Zaikov, G. (eds), *Polymer nanocomposites recent advances* (Nova Science, New York, 2008).
22. Paul, D. R. & Robeson, L. M. Polymer technology: Nanocomposites. *Polymer* **49**, 3187–3204 (2008).
23. Boatner, L. & Sales, B. inventors; The United States of America as represented by the Department of Energy, assignee. Lead iron phosphate glass as a containment medium for disposal of high-level nuclear waste. United States Patent US 4,847,008. 1989 Jul 11.
24. Quinn, C. J., Frayer, P. D. & Beall, G. H. in *Polymeric Materials Encyclopedia* (Vol. 4) [Salamone, J. C. (ed.)], [2766–2777] (CRC Press, Florida, 1996).
25. Brow, R. K. in *Structure, Properties and Applications of Phosphate and Phosphate-containing Glasses* [Brow, R. K. (ed.)] [1–28] (Elsevier, North-Holland, 2000).
26. Einstein, A. in *Reader's Digest* October, 1977).
27. Sanford, L. & Tick, P. inventors; Corning Glass Works, assignee. Tin-phosphorus oxyfluoride glasses. United States Patent US 4,314,031. 1982 Feb 2.
28. Beall, G. & Quinn, C. inventors; Corning, Inc., assignee. Alkali zinc halophosphate glasses. United States patent US 5,071,795. 1991 Dec 10.
29. Beall, G. & Pierson, J. inventors; Corning, Inc., assignee. Zinc sulfophosphate glasses. United States patent US 5,328,874. 1994 Jul 12.
30. McKenna, G. B. & Simon, S. L. in *Handbook of Thermal Analysis and Calorimetry, Applications to Polymers and Plastics* [Vol. 3] [Cheng, S. Z. D. (ed.)] [49–110] (Elsevier, North-Holland, 2002).

Acknowledgments

J.U.O. gratefully acknowledges the U.S. Department of State, the French Ministry of Higher Education and Research, and the Franco-American Commission for jointly funding his *Fulbright-Tocqueville Distinguished Chair award in Engineering* at the University of Lyon 1; and IMP CNRS 5223 colleagues for their hospitality during his temporary residence in Lyon. This work was partially supported by the U.S. National Science Foundation Office of International and Integrative Activities through IIA-1346898 grant award and of Division of Materials Research through DMR-1360006, and by Chemtura Corporation.

Author contributions

R.F. and J.U.O. conceived the research and wrote the paper, I.B. prepared the glass samples, R.F. and V.B.-L. performed the measurements, and all authors analysed and discussed the results.

Additional information

Supplementary information accompanies this paper at <http://www.nature.com/scientificreports>

Competing financial interests: The authors declare no competing financial interests.

How to cite this article: Fulchiron, R., Belyamani, I., Otaigbe, J.U. & Bounor-Legaré, V. A simple method for tuning the glass transition process in inorganic phosphate glasses. *Sci. Rep.* **5**, 8369; DOI:10.1038/srep08369 (2015).



This work is licensed under a Creative Commons Attribution 4.0 International License. The images or other third party material in this article are included in the article's Creative Commons license, unless indicated otherwise in the credit line; if the material is not included under the Creative Commons license, users will need to obtain permission from the license holder in order to reproduce the material. To view a copy of this license, visit <http://creativecommons.org/licenses/by/4.0/>